

Name _____

Date _____

Atomic Structure Study Guide

Atomic Structure

1. Which subatomic particle has a positive charge? _____
2. Which subatomic particle has a neutral charge? _____
3. Which subatomic particle has a negative charge? _____
4. Which two subatomic particles are found inside the nucleus? _____

5. Which subatomic particle orbits the nucleus? _____

Periodic Table

6. How are atoms arranged on the periodic table? _____

7. How do you find the number of protons in an element? _____
8. How do you find the number of electrons in an element? _____
9. How do you find the mass number of an element? _____

10. How do you find the number of neutrons in an element? _____

11. Where are metals located on the periodic table? _____
12. Where are non-metals located on the periodic table? _____
13. Where are metalloids located on the periodic table? _____

Isotopes

1. What makes an isotope of an element different from another isotope of the same element?

2. Which particle is the same as the atomic number for any element?

3. How can you find the number of neutrons in the most abundant isotope for any element?

4. How many neutrons are in Carbon-12?

5. How many neutrons are in Carbon-14?

6. How many protons are in Calcium - 40?

7. How many electrons are in Calcium – 40 ?

8. How many neutrons are in Calcium – 40 ?

9. Draw the isotope symbol for Calcium-40:

10. How many protons are in Calcium – 42?

11. How many neutrons are in Calcium – 42?

12. Draw the isotope symbol for Calcium-42:

13. Which is the most abundant Calcium isotope?

14. What does it mean when it's said that many isotopes are radioactive?

15. List two common uses for radioactive isotopes:

16.

17.
