

Notes for Acids and Bases

- **pH** – stands for “percent Hydrogen” and is a measurement of how many H^+ ions a chemical releases when dissolved in water.
 - Scale – 0 – 14
 - Each change in a number on the pH scale represents a change of 10%. Substance with a pH of 3 is 10X more acidic than pH of 4 and 100X more acidic than pH of 5.
- **Acids** – Release lots of H^+ ions and have pH below 7.
 - Taste sour and are found in a lot of our foods.
 - Stronger acids have lower numbers.
- **Bases** – (alkaline) Most contain an hydroxide group (OH) that removes H^+ from a solution to form H_2O .
 - pH above 7 with stronger bases having higher numbers.
- **Neutral** – pH of 7 – distilled water

- **Corrosive** – can burn skin, both acids and bases can be corrosive.
- **Neutralization** – when an acid and a bases are combined, the pH changes to neutral and salt and water are formed.
- **pH Buffers** – chemicals that help substances resist large changes in pH.
- **pH indicators** – measure pH.
- **Litmus papers** – tell if a substance is an acid or a base.
 - Pink – acid
 - Blue - base