Notes for Electron Arrangement

- Electron orbit the nucleus in energy levels
  - 1st energy level (closest to nucleus can hold 2 electrons)
  - 2nd energy level - 8 electrons and 3rd energy level – 18 electrons
- The rows on the periodic table correspond to the number of energy levels (called periods).
  - H and He – 1st row – 1 energy level; Li → Ne 2nd row & 2 energy levels
- Valence electrons are electrons in the outermost energy level
  - These determine how elements behave in a chemical reaction
  - Elements with same number of valence electrons have similar properties and will behave the same way during chemical reactions
- The columns on the periodic table correspond to the number of valence electrons (called groups)